

การสังเคราะห์ผลลัพธ์โดยผลลัพธ์ของเครื่องมือที่มีรูปรุนขนาดนาโน

นางสาวมยุรี ศรีนันทากุล

# ศูนย์วิทยทรัพยากร

วิทยานิพนธ์นี้เป็นส่วนหนึ่งของการศึกษาตามหลักสูตรปริญญาวิทยาศาสตร์มหาบัณฑิต  
สาขาวิชาปีตรีเคมีและวิทยาศาสตร์พลีเมอร์  
คณะวิทยาศาสตร์ จุฬาลงกรณ์มหาวิทยาลัย

ปีการศึกษา 2547

ISBN: 974-17-6707-2

ลิขสิทธิ์ของจุฬาลงกรณ์มหาวิทยาลัย

SYNTHESIS OF POLYMER BRUSH BY NANOPOROUS  
SURFACE-INITIATED POLYMERIZATION

Miss Mayuree Srinunthakul

A Thesis Submitted in Partial Fulfillment of the Requirements  
for the Degree of Master of Science in Petrochemistry and Polymer Science

Faculty of Science

Chulalongkorn University

Academic Year 2004

ISBN: 974-17-6707-2

**Thesis Title** Synthesis of Polymer Brush by Nanoporous Surface-Initiated Polymerization  
**By** Miss Mayuree Srinunthakul  
**Field of Study** Petrochemistry and Polymer Science  
**Thesis Advisor** Assistant Professor Vipavee P. Hoven, Ph.D.

Accepted by the Faculty of Science, Chulalongkorn University in Partial  
Fulfillment of the Requirements for the Master's Degree

 ..... Dean of the Faculty of Science  
(Professor Piamsak Menasveta, Ph.D.)

## Thesis committee

 ..... Chairman  
(Professor Pattarapan Prasassarakich, Ph.D.)

*Vipavee Hoven* Thesis Advisor  
(Assistant Professor Vipavee P. Hoven, Ph.D.)

*Suda Kiatkamjornwong* Member  
(Professor Suda Kiatkamjornwong, Ph.D.)

Warinthorn Chavasiri Member  
(Assistant Professor Warinthorn Chavasiri, Ph.D.)

มยุรี ศรีนันทากุล: การสังเคราะห์พอลิเมอร์บัชโดยพอลิเมอไรเซ็นต์ริเริ่มบนพื้นผิวที่มีรูพรุนขนาดนาโน (SYNTHESIS OF POLYMER BRUSH BY NANOPOROUS SURFACE-INITIATED POLYMERIZATION) อาจารย์ที่ปรึกษา: ผศ.ดร.วิภาวดี ไฮเว่น; 125 หน้า ISBN: 974-17-6707-2

ใช้ชับสเตรตซิลิกอนออกไซด์กราฟต์ด้วยโมเลกุลชั้นเดียวของทริสไตรเมทิลไซโลกซิไซลิล (ทริสทีเอ็มเอส) เป็นแม่แบบระดับนาโนเมตรในการควบคุมความหนาแน่นการกราฟต์ของพอลิเมอร์บัช การสังเคราะห์พอลิเมอร์บัชโดยการริเริ่มปฏิกิริยาพอลิเมอไรเซ็นต์ของ 2-เมทอกิซิโอล อิลออกซิเอทธิลฟอลโซฟิลโคลีน (เอ็มพีซี) และ เทอร์ต-บิวิวิเมทาคริเลต (ที-บีเอ็มเอก) จากหมู่แอลฟ้าไบโรเมอสเทอเรชีนติดกับหมู่ไซลานอลที่เหลืออยู่บนพื้นผิวหลังจากถูกปักคลุมด้วยทริสทีเอ็มเอสบริมาณต่างๆ ด้วยอะตอมกรานส์เฟอร์แรดิคัลพอลิเมอไรเซ็นต์ โดยใช้คอปเปอร์ไบรอนด์/ไบไฟริดิน และคอปเปอร์ไบรอนด์/พีเอ็มดีอีทีเอเป็นระบบเร่งปฏิกิริยาในการสังเคราะห์พีเอ็มพีซี และพีทีบีเอ็มเอก ตามลำดับ ร้อยละการปักคลุมของทริสทีเอ็มเอสสั่งผลกระทบอย่างมีนัยสำคัญต่อความหนาและสัณฐานวิทยาของพอลิเมอร์บัช ส่วนที่ยื่นขึ้นชี้สังเกตเห็นได้จากการภาพเออฟเอ็มของพีเอ็มพีซีบัชเป็นหลักฐานที่แสดงให้เห็นว่า พีเอ็มพีซีบัชมีการกระจายตัวในระดับนาโนเมตรบนชับสเตรตขนาดของส่วนที่ยื่นขึ้นและความชุรุวะของพื้นผิวมีลักษณะสอดคล้องเป็นอย่างดีกับความหนาแน่นการกราฟต์ของพีเอ็มพีซีบัช การที่พีทีบีเอ็มเบอร์รัชที่เติบโตจากชับสเตรตที่มีรูขนาดนาโนเมตรแทบจะไม่ปรากฏลักษณะสำคัญใดๆ แสดงว่าการจับตัวกันเองเป็นกลุ่มก้อนของพีเอ็มพีซีบัชเกิดจากการเข้ากันไม่ได้ระหว่างเฟสของพีเอ็มพีซีบัชที่มีสมบัติชอบน้ำ และหมู่ทริสทีเอ็มเอสที่มีสมบัติไม่ชอบน้ำ

สาขาวิชา ปิโตรเคมีและวิทยาศาสตร์พอลิเมอร์ ลายมือชื่อนิสิต นางสาวนันดาภรณ์  
ปีการศึกษา 2547 ลายมือชื่ออาจารย์ที่ปรึกษา \_\_\_\_\_

ลายมือชื่ออาจารย์ที่ปรึกษาร่วม \_\_\_\_\_

# # 4573826723: MAJOR PETROCHEMISTRY AND POLYMER SCIENCE

KEYWORD: POLYMER BRUSH/ PHOSPHOLIPID POLYMER/ SURFACE-INITIATED POLYMERIZATION/ ATOM TRANSFER RADICAL POLYMERIZATION/ NANO-SCALE TEMPLATE

MAYUREE SRINUNTHAKUL: SYNTHESIS OF POLYMER BRUSH BY NANOPOROUS SURFACE-INITIATED POLYMERIZATION. THESIS  
ADVISOR: VIPAVEE P. HOVEN, Ph.D. 125 pp ISBN: 974-17-6707-2

Chemically grafted tris(trimethylsiloxy)silyl (tris(TMS)) monolayer on the silicon oxide substrate was used as a nanometer-scale template for controlling the graft density of polymer brushes. Polymer brushes were synthesized by surface-initiated polymerization of 2-methacryloyloxyethyl phosphorylcholine (MPC) and *tert*-butyl methacrylate (Pt-BMA) from  $\alpha$ -bromoester groups tethered to the residual silanol groups on the surface after creating a range of tris(TMS) coverage *via* atom transfer radical polymerization (ATRP). CuBr/bpy or CuBr/PMDTA was used as a catalytic system for PMPC and Pt-BMA synthesis, respectively. The percentage of tris(TMS) coverage significantly influenced the thickness and morphology of polymer brushes. Protrusions representing self-aggregation of PMPC brushes in nanopores as visualized by AFM analysis evidently suggested that PMPC brushes distributed nanoscopically on the surface. The size of protrusion and surface roughness corresponded quite well with the graft density of PMPC brushes. The fact that Pt-BMA brushes grown from nanopores were almost featureless implies that self-aggregation of PMPC brushes is truly a consequence of phase incompatibility between hydrophilic PMPC brushes and hydrophobic tris(TMS).

Field of study Petrochemistry and Polymer Science Student's signature Mayuree Srinunthakul  
Academic year ..... 2004 Advisor's signature V. Hoven  
Co-advisor's signature .....

## ACKNOWLEDGEMENTS

I would like to express my heartfelt gratitude and appreciation to my advisor, Asst. Prof. Dr. Vipavee P. Hoven, for supporting me both in science and in life, and encouraging me throughout the course of my study. I am sincerely grateful to the members of the thesis committee, Prof. Dr. Pattarapan Prasassarakich, Prof. Dr. Suda Kiatkamjornwong, Asst. Prof. Dr. Warinthorn Chavasiri for their comments, suggestions and time to read the thesis.

Many thanks go to Prof. Dr. Kazuhiko Ishihara for his support in MPC monomer, Asst. Prof. Dr. Junji Watanabe for GPC analysis, Assoc. Prof. Yasuhiko Iwasaki for his support in materials, XPS and GPC analyses as well as fruitful suggestions on solving the research problems.

Special thanks go to Electronic Research Center, Faculty of Engineering, King Mongkut's Institute of Technology, Ladkrabang for ellipsometry facility, National Metal and Materials Technology Center for contact angle goniometer and Capability Building Unit in Nanoscience and Nanotechnology, Department of Physics, Faculty of Science, Mahidol University for AFM facility. A grateful acknowledgment to the Thailand Research Fund, the research funding for graduate students from Ministry of University Affairs, the Teaching Assistant Fund from Chulalongkorn University and Research Assistant Fund from Organic Synthesis Research Unit (OSRU) for financial support.

Many thanks go to all OSRU members for their assistance, suggestions concerning experimental techniques and their kind helps during my thesis work.

Finally, I would like to especially thank my family members: father, mother, my elder sister, two elder brothers and relatives for their love, kindness and support throughout my entire study.

## CONTENTS

## **CONTENTS (Continued)**

	Page	
3.2.1	Ellipsometry.....	40
3.2.2	X-ray Photoelectron Spectroscopy (XPS).....	40
3.2.3	Nuclear Magnetic Resonance Spectroscopy (NMR).....	40
3.2.4	Contact Angle Measurement.....	41
3.2.5	Atomic Force Microscopy (AFM).....	41
3.3	Synthesis of $\alpha$ -Bromoester Derivatives to be used as Initiators.....	41
3.3.1	Synthesis of Vinyl-terminated $\alpha$ - Bromoisobutyrate Compounds.....	41
3.3.2	Synthesis of Silane Compounds by Hydrosilylation of Vinyl-terminated $\alpha$ - Bromoisobutyrate Compounds.....	42
3.3.3	Synthesis of Prop-2-bromo-2-methyl- propionate as a “Sacrificial” Initiator.....	43
3.4	Pretreatment of Silicon Substrates.....	44
3.5	Preparation of Silicon-supported Mixed Tris(TMS) / silanol Monolayer.....	44
3.6	Preparation of Silicon-supported $\alpha$ -Bromoisobuty- rate Monolayer and Silicon-supported Mixed Tris(TMS)/ $\alpha$ -Bromoisobutyrate Monolayer.....	45
3.7	Preparation of Polymer Brushes.....	46
3.7.1	Surface-initiated Polymerization of 2- methacryloyloxyethyl phosphorylcho- line) (MPC).....	46
3.7.2	Surface-initiated Polymerization of <i>tert</i> - Butyl Methacrylate ( <i>t</i> -BMA).....	49

## CONTENTS (Continued)

	Page
CHAPTER IV : RESULTS AND DISCUSSION.....	52
4.1    Synthesis of $\alpha$ -Bromoester Derivatives as Initiators.....	52
4.1.1    Synthesis of Vinyl-terminated $\alpha$ -Bromo isobutyrate Compounds.....	53
4.1.2    Synthesis of Silane Compounds by Hydrosilylation of Vinyl-terminated $\alpha$ - Bromoisobutyrate Compounds.....	54
4.1.3    Synthesis of Propyl (2-bromo-2-methyl) propionate as a “Sacrificial” Initiator.....	60
4.2    Preparation of Silicon -supported Mixed Tris(TMS) / silanol Monolayer.....	61
4.3    Preparation of Silicon-supported $\alpha$ -Bromoisobutyrate Monolayer and Silicon-supported Mixed Tris(TMS)/ $\alpha$ - Bromoisobutyrate Monolayer.....	65
4.4    Preparation of Polymer Brushes.....	68
4.4.1    Surface-initiated Polymerization of MPC from Silicon-supported $\alpha$ -Bromo isobutyrate Monolayer.....	68
4.4.2    Surface-initiated Polymerization of <i>t</i> -BMA from Silicon-supported $\alpha$ -Bromo isobutyrate Monolayer.....	79
4.4.3    Surface-initiated Polymerization of MPC from Silicon-supported Mixed Tris(TMS)/ $\alpha$ -Bromoisobutyrate Monolayer.....	84
4.4.4    Surface-initiated Polymerization of <i>t</i> -BMA from Silicon-supported Mixed Tris(TMS)/ $\alpha$ -Bromoisobutyrate Monolayer.....	90

## CONTENTS (Continued)

	Page
4.5    Surface Topography of Polymer Brushes.....	92
CHAPTER V : CONCLUSIONS .....	102
REFERENCES.....	104
APPENDICES.....	115
APPENDIX A.....	116
APPENDIX B.....	120
VITAE.....	125


  
 ศูนย์วิทยบรังษยการ  
 จุฬาลงกรณ์มหาวิทยาลัย

## LIST OF FIGURES

Figure	Page
2.1 Strategy for amplification of a patterned SAM prepared by microcontact printing into a patterned polymer brush.....	6
2.2 AFM images (height/phase) of 11K PS-COOH adsorbed from toluene to tris(TMS) modified surfaces.....	8
2.3 Molecular weight conversion curves for various kinds of polymerization methods: (A) living polymerization; (B) free radical polymerization; and (C) condensation polymerization.....	10
2.4 Architectural forms of polymers available by living polymerization techniques.....	11
2.5 The mechanism of ATRP.....	12
2.6 Equilibrium reaction in ATRP.....	13
2.7 The rotation of the bpy ligands from the tetrahedral and co-ordination of halide at the Cu center.....	14
2.8 Complex formation equilibrium in polar and nonpolar solvents.....	15
2.9 Examples of polymer systems comprising polymer brushes.....	18
2.10 Classification of linear polymer brushes, (a <sub>1</sub> -a <sub>4</sub> ) homopolymer brushes; (b) mixed homopolymer brush; (c) random copolymer brush; (d) block copolymer brush.....	20
2.11 Preparation of polymer brushes by “physisorption”, “grafting to” and “grafting from”.....	22
2.12 AFM images of PMMA brushes grown from gold evaporated onto mica after 50 min of reaction time: (a) 10% of active initiator, (b) 25% of active initiator and (c) 100% of active initiator.....	25
2.13 AFM images of (a) Si-H surface, (b) Si- R <sub>3</sub> Br surface, and (c) Si-g-PMMA surface.....	26
2.14 Schematic of the geometry of an ellipsometry experiment.....	28
2.15 Schematic representation of the Young’s equation .....	28
2.16 Schematic representation of wettability.....	29

## LIST OF FIGURES (Continued)

Figure	Page
2.17 Schematic diagram of the X-ray photoelectron emission process.....	30
2.18 General schematic drawing of the XPS instrument.....	31
2.19 Schematic diagram of an atomic force microscope.....	32
2.20 Schematic representation of the gel permeation chromatography.....	37
4.1 Mechanism of nucleophilic acyl substitution of vinyl-terminated alcohol with 2-bromoisobutyryl bromide.....	53
4.2 Mechanism of hydrosilylation using chloroplatinic acid as a catalyst....	55
4.3 $^1\text{H}$ NMR spectra of (A) Prop-2-enyl (2-bromo-2-methyl) propionate and (B) 3-(Dimethylethoxysilyl) propyl (2-bromo-2-methyl) propionate.....	57
4.4 $^1\text{H}$ NMR spectra of (A) Hex-5-enyl (2-bromo-2-methyl) propionate and (B) 3-(Dimethylethoxysilyl) hexyl (2-bromo-2-methyl) propionate....	58
4.5 $^1\text{H}$ NMR spectra of (A) Dec-5-enyl (2-bromo-2-methyl) propionate and (B) 3-(Dimethylethoxysilyl) decyl (2-bromo-2-methyl) propionate.....	59
4.6 $^1\text{H}$ NMR spectrum of Propyl(2-bromo-2-methyl)propionate .....	60
4.7 Ellipsometric thickness of tris(TMS) monolayer as a function of reaction time.....	62
4.8 Water contact angle of tris(TMS) monolayer as a function of reaction time.....	62
4.9 Schematic representation of nanopores in tris(TMS) monolayer.....	64
4.10 The activation/deactivation cycles of ATRP process.....	69
4.11 PMPC thickness versus polymerization time in the presence of “added” initiator: [MPC] : [added initiator] = 200:1 (○) and [MPC] : [added initiator] = 50:1(●) using methanol/water = 4 : 1 (v/v) as a solvent.....	70

## LIST OF FIGURES (Continued)

Figure	Page
4.12 Ellipsometric thickness of PMPC brushes versus polymerization time for targeted DP = 200 in methanol:H <sub>2</sub> O = 4:1 (v/v) (■) and pure methanol(□).....	71
4.13 Water contact angle data of PMPC brushes versus polymerization time for targeted DP = 200 using methanol/water = 4 : 1 (v/v) as a solvent: $\theta_a$ (●) and $\theta_r$ (○).....	72
4.14 Ellipsometric thickness (●) and receding water contact angle ( $\theta_r$ ) (○) of PMPC brushes versus polymerization time for targeted DP = 200 using methanol:water = 4:1 (v/v) as a solvent.....	73
4.15 Ellipsometric thickness (■) and receding water contact angle ( $\theta_r$ ) (□) of PMPC brushes versus polymerization time for targeted DP = 200 using methanol as a solvent.....	73
4.16 Ellipsometric thickness of PMPC brushes versus polymerization time for different [CuBr]/[CuBr <sub>2</sub> ] ratios at targeted DP of 200 using methanol:water = 4:1 (v/v) as a solvent.....	75
4.17 The molecular weight ( $\overline{M}_w$ ) (●) and molecular weight distribution $(\overline{M}_w/\overline{M}_n)$ (○) of free PMPC for targeted DP = 200 produced in methanol:water = 4:1 (v/v) as a function of polymerization time.....	76
4.18 Relationship between the thickness (●) and graft density (■) of PMPC brushes with the molecular weight ( $\overline{M}_n$ ) of free PMPC for targeted DP = 200 produced in methanol:water = 4:1 (v/v).....	77

## LIST OF FIGURES (Continued)

Figure	Page
4.19 Relationship between the thickness of PMPC brushes with the molecular weight ( $\overline{M}_n$ ) of free PMPC for targeted DP = 50 produced in methanol:water = 4:1 (v/v).....	77
4.20 Ellipsometric thickness of PtBMA brushes versus polymerization time..	81
4.21 Water contact angle data of PtBMA brushes versus polymerization time for targeted DP = 200.....	81
4.22 Relationship between the thickness of Pt-BMA brushes with the molecular weight ( $\overline{M}_n$ ) of free PMPC for targeted DP = 200 produced in toluene at 90 °C.....	82
4.23 The molecular weight ( $\overline{M}_n$ ) (●) and molecular weight distribution ( $\overline{M}_w/\overline{M}_n$ ) (○) of free PtBMA produced in toluene at 90°C as a function of polymerization time.....	83
4.24 Ellipsometric thickness of PMPC brushes grown from silicon-supported mixed tris(TMS)/ $\alpha$ -bromoisobutyrate monolayer using methanol (●) or water (○) as a solvent.....	85
4.25 Water contact angle of PMPC brushes grown from silicon-supported mixed tris(TMS)/ $\alpha$ -bromoisobutyrate monolayer using methanol as a solvent for 5 h.....	86
4.26 Ellipsometric thickness (○) and receding water contact angle (●) of the silicon surface having 82% tris(TMS)/PMPC brushes as a function of polymerization time.....	87

## LIST OF FIGURES (Continued)

Figure	Page
4.27 Ellipsometric thickness of PMPC brushes grown from silicon-supported mixed tris(TMS)/ $\alpha$ -bromoisobutyrate monolayer having 82% tris(TMS) coverage versus grafting time of initiator in methanol for 5 h: n <sub>3</sub> (●), n <sub>6</sub> (○)and n <sub>10</sub> (▲). ....	88
4.28 Ellipsometric thickness of PMPC brushes grown from silicon-supported mixed tris(TMS)/ $\alpha$ -bromoisobutyrate monolayer having 82% tris(TMS) coverage versus polymerization time in methanol : n <sub>3</sub> (●), n <sub>6</sub> (○)and n <sub>10</sub> (▲).....	89
4.29 The thickness of Pt-BMA brushes grown from silicon-supported mixed tris(TMS)/ $\alpha$ -bromoisobutyrate monolayer using toluene as a solvent for 5 h polymerization time.....	91
4.30 Water contact angle of Pt-BMA brushes grown from silicon-supported mixed tris(TMS)/ $\alpha$ -bromoisobutyrate monolayer using toluene as a solvent for 5 h polymerization time.....	91
4.31 AFM images of silicon-supported mixed tris(TMS)/silanol monolayer having varied %tris(TMS) coverage: (a) 0 %, (b) 51%, (c) 66 %, and (d) 82% and silicon-supported mixed tris(TMS)/ $\alpha$ -bromoisobutyrate monolayer having varied %tris(TMS) coverage: (e) 0 %, (f) 51%, (g) 66 %, and (h) 82 %.....	93

## LIST OF FIGURES (Continued)

Figure	Page
4.32 AFM images of silicon-supported mixed tris(TMS)/PMPC brushes prepared in methanol for 1 h having varied having varied %tris(TMS) coverage : (a) 0%, (b) 51%, (c) 66% and (d) 82%.....	95
4.33 AFM images of silicon-supported mixed tris(TMS)/PMPC brushes having 82% tris(TMS) coverage prepared in methanol for 1 h by controlling grafting time of initiator: (a) 1 day, (b) 2 days, (c) 3 days and (d) 4 days.....	97
4.34 Schematic representation of possible orientation of polymer brushes grown from nanopores having different graft densities.....	98
4.35 AFM images of silicon-supported mixed tris(TMS)/PMPC brushes prepared in methanol for 1 h using grafting time of initiator for 1 day and having varied % tris(TMS) coverage: (a) 0%, (b) 51%, (c) 66% and (d) 82%.....	99
4.36 AFM images of silicon-supported mixed tris(TMS)/Pt-BMA brushes having 82% tris(TMS) coverage prepared in toluene at 90 °C for 1 h by controlling grafting time of initiator: (a) 1 day and (b) 4 days.....	100
4.37 AFM images of silicon-supported mixed tris(TMS)/Pt-BMA brushes having 82% tris(TMS) coverage prepared in toluene at 90 °C for 5 h by controlling grafting time of initiator: (a) 1 day and (b) 4 days.....	101

## LIST OF TABLES

Table	Page
4.1 Characteristic $^1\text{H}$ NMR peaks and %yield of vinyl-terminated $\alpha$ -bromo isobutyrate compounds.....	54
4.2 Characteristic $^1\text{H}$ NMR peaks and %yield of silane compounds.....	56
4.3 XPS data ( $15^\circ$ take-off angle), water contact angle and calculated % coverage of tris(TMS) monolayer as a function of reaction time.....	64
4.4 XPS atomic composition ( $15^\circ$ takeoff angle) and contact angle data for silicon-supported mixed tris(TMS)/ $\alpha$ -bromo isobutyrate ( $n = 3$ ) monolayer using 4 days of reaction.....	67
4.5 XPS elemental surface composition (%) of silicon surfaces before and after the formation of PMPC brushes.....	79
4.6 Average roughness of silicon-supported mixed tris(TMS)/silanol monolayer determined by AFM analysis.....	92
4.7 Average roughness of silicon-supported mixed tris(TMS)/ PMPC brushes prepared in methanol determined by AFM analysis.....	96
A-1 Average thickness of PMPC brushes versus polymerization time in the presence of “added” initiator.....	117
A-2 Average thickness of PMPC brushes versus polymerization time for different $[\text{CuBr}]/[\text{CuBr}_2]$ ratios at targeted DP of 200 using methanol:water = 4:1 (v/v) as a solvent.....	117
A-3 Average thickness, water contact angle, molecular weight ( $\overline{M}_n$ ) and molecular weight distribution ( $\overline{M}_w/\overline{M}_n$ ) of free PMPC for targeted DP = 50 produced in methanol:water = 4:1 (v/v) as a function of polymerization time.....	118

## LIST OF TABLES (Continued)

	Page
Table	
A-4 Average thickness, water contact angle, molecular weight ( $\overline{M}_n$ ) and molecular weight distribution ( $\overline{M}_w/\overline{M}_n$ ) of free PMPC for targeted DP = 200 produced in methanol:water = 4:1 (v/v) as a function of polymerization time.....	118
A-5 Average thickness, water contact angle, molecular weight ( $\overline{M}_n$ ) and molecular weight distribution ( $\overline{M}_w/\overline{M}_n$ ) of free Pt-BMA for targeted DP = 200 produced in toluene at 90°C as a function of polymerization time. ....	119
B-1 Average thickness of tris(TMS) monolayer as a function of reaction time calculated from ellipsometric data and advancing and receding water contact angle as a function of time (days).....	121
B-2 Average thickness of silicon-supported mixed tris(TMS)/ $\alpha$ -bromoiso- butyrate (n = 3) monolayer using 1 - 4 days of reaction.....	121
B-3 Average advancing and receding water contact angle of silicon- supported mixed tris(TMS)/ $\alpha$ -bromoisoobutyrate (n = 3) monolayer using 1 - 4 days of reaction.....	122
B-4 Average thickness of PMPC brushes grown from silicon-supported mixed tris(TMS)/ $\alpha$ -bromoisoobutyrate monolayer using methanol or water.....	122
B-5 Average thickness of PMPC brushes grown from silicon-supported mixed tris(TMS)/ $\alpha$ -bromoisoobutyrate monolayer having 82% tris(TMS) coverage versus grafting time of initiator in methanol for 5 h.....	123

## LIST OF TABLES (Continued)

Table	Page
B-6 Average thickness of PMPC brushes grown from silicon-supported mixed tris(TMS)/ $\alpha$ -bromo isobutyrate monolayer having 82%tris(TMS) coverage ( $I = 4$ days) versus polymerization time in methanol.....	123
B-7 Average thickness of Pt-BMA brushes grown from silicon-supported mixed tris(TMS)/ $\alpha$ -bromo isobutyrate monolayer using toluene as a solvent for 5 h calculated from ellipsometric data and advancing and receding water contact angle as a function of time.....	124

## LIST OF ABBREVIATIONS

AFM	: Atomic force microscopy
ATRP	: Atom transfer radical polymerization
Å	: Angström
<i>t</i> -BMA	: <i>tert</i> -Butyl methacrylate
bpy	: 2,2'-Bipyridyl
CDCl <sub>3</sub>	: Deuterochloroform
CuBr	: Copper (I) bromide
CuBr <sub>2</sub>	: Copper (II) bromide
°C	: Degree Celsius
d	: doublet (NMR)
D	: Polydispersity
DEA	: 2-(Diethylamino)ethyl methacrylate
D <sub>2</sub> O	: Deuterium oxide
<i>dp</i>	: Degree of polymerization
Eq.	: Equation
GPC	: Gel permeation chromatography
<i>k</i> <sub>act</sub>	: The activation rate parameter
<i>k</i> <sub>deact</sub>	: The deactivation rate parameter

m	: multiplet (NMR)
MeOH	: Methanol
MEO2PC	: 2-methacryloyloxy ethoxyethyl phosphorylcholine
mg	: miligram
MgSO <sub>4</sub>	: Magnesium sulfate
min	: minute
mL	: mililiter
mm	: milimeter
mM	: milimolar
MMA	: Methyl methacrylate
mmol	: milimole
$\overline{M}_n$	: Number average molecular weight
MPC	: 2-Methacryloyloxyethyl phosphorylcholine
$\overline{M}_w$	: Weight average molecular weight
nm	: nano meter
NMR	: Nuclear magnetic resonance spectroscopy
PDI	: Polydispersity Index
PMMA	: Poly(methyl methacrylate)
PMPC	: Poly(2-methacryloyloxyethyl phosphorylcholine)
Pt-BMA	: Poly( <i>tert</i> -butyl methacrylate)

ppm	: part per million
q	: quartet (NMR)
s	: singlet (NMR)
SAM	: Self-assembled monolayer
S.D.	: Standard deviation
SEM	: Scanning electron microscopy
t	: triplet (NMR)
THF	: Tetrahydrofuran
tris(TMS)Cl	: tris (Trimethylsiloxy)chlorosilane
tris(TMS)	: tris (Trimethylsiloxy)silyl
$\mu\text{L}$	: microliter
wt%	: percentage of weight
XPS	: X-ray photoelectron spectroscopy

ศูนย์วิทยาหัตถกรรม  
จุฬาลงกรณ์มหาวิทยาลัย